**Compounds Quiz REVIEW**

*This quiz is out of 15 marks. You may use your periodic table for this quiz.*

***Vocabulary***

*Ionic bond covalent bond diatomic molecule Bohr model valence shell valence electron*

1. Answer the following questions for Aluminum.
2. Draw a Bohr diagram for Aluminum.
3. How many shells does it have?
4. How many electrons are in its valence shell?
5. In a reaction, does it transfer, share or take electrons?
6. Write the chemical formulae for the following compounds:
7. Strontium phosphorus
8. Aluminum bromide
9. Silver nitride
10. Zinc iodide
11. Cesium selenide
12. Write the formulae for the following compounds:

a.) dimolybdenum tetraiodide

b.) vanadium tribromide

c.) tribromine octoxide

d.) disilicon hexaiodide

e.) tungsten hexafluoride

1. Name the following compounds:
2. CaCl2 b.)MgO c.) Ca3P2 d.) P4S5 e.) SeF6
3. Name the following compounds:
4. PF3 b.)SCl2 c.) CCl4 d.) N2O e.) P2O3
5. How many atoms are in a molecule of -

a.) NaCl b.) CaCl2 C.) Al2S3

1. Define covalent bond, explaining what happens to the electrons.
2. a.) What is a diatomic molecule?
3. Which elements form diatomic molecules?
4. What patterns are witnessed on the periodic table. Describe three (in relation to shells, electrons and increasing numbers)